Common Polyatomic Ions

Name	Ion Formula	Name	Ion Formula
ammonium	NH_4^+	hydroxide	OH^-
acetate[1]	$C_2H_3O_2^-$	iodate	IO_3^-
hypobromite	BrO^-	nitrite	NO_2^-
bromite	BrO_2^-	nitrate	NO_3^-
bromate	BrO_3^-	oxalate	$C_2O_4^{2-}$
perbromate	BrO_4^-	permanganate	MnO_4^-
carbonate	CO_3^{2-}	peroxide	O_2^{2-}
hydrogen carbonate	HCO_3^-	phosphate	PO_4^{3-}
hypochlorite[2]	ClO^-	hydrogen phosphate	HPO_4^{2-}
chlorite	ClO_2^-	dihydrogen phosphate	$H_2PO_4^-$
chlorate	ClO_3^-	phosphite	PO_3^{2-}
perchlorate	ClO_4^-	selenate	SeO_4^{2-}
chromate	CrO_4^{2-}	sulfate	SO_4^{2-}
dichromate	$Cr_2O_7^{2-}$	hydrogen sulfate	HSO_4^-
cyanide	CN^-	thiosulfate	$S_2O_3^{2-}$
cyanate[3]	CNO^-	sulfite	SO_3^{2-}
thiocyanate	SCN^-	hydrogen sulfite	HSO_3^-

The word “ion” is assumed for each of the polyatomic ion names in this table.

[1] The formula for the acetate ion can also be written CH_3COO^- and $CH_3CO_2^-$

[2] The formula for the hypochlorite can also be written OCl^-

[3] The formula for the cyanate ion can also be written OCN^-

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